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Composition Worksheet - Solutions Find  
the percent compositions of all of the  
elements in the following compounds: 1)  
CuBr<sub>2</sub> Copper (II) Bromide Cu: 28.45%  
Br: 71.55% 2) NaOH Sodium Hydroxide  
Na: 57.48% O: 39.99% H: 2.53% 3) (NH<sub>4</sub>)<sub>2</sub>S Ammonium

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Chemists often need to know what elements are present in a compound and in what percentage. The percent composition is the percent by mass of each element in a compound. It is calculated in a similar way that we just indicated for the peanut butter. 
$$\text{\% by mass} = \frac{\text{mass}}$$

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compositions of all of the elements in  
the following compounds: 1) CuBr 2  
Copper (II) Bromide Cu: 28.45% Br:  
71.55% 2) NaOH Sodium Hydroxide Na:  
57.48% O: 39.99% H: 2.53% 3) (NH<sub>4</sub>)<sub>2</sub>S  
Ammonium Sulfide N: 41.1% H: 11.8% S:  
47.1% 4) N<sub>2</sub>S<sub>2</sub>

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COMPLETE percentage composition (by  
mass) of the following: (a) cuprous oxide  
 $2 \text{ mol Cu} \times 63.55 \text{ g Cu} = 127.10$

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Copper (II) Bromide Cu: 28.45% Br:  
71.55% 2) NaOH Sodium Hydroxide Na:  
57.48% O: 39.99% H: 2.53% 3) (NH<sub>4</sub>)<sub>2</sub>S  
Ammonium Sulfide N: 41.1% H: 11.8% S:  
47.1% 4) N<sub>2</sub>S<sub>2</sub>

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