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The formula is: $M_1 V_1 = M_2 V_2$ Where M_1 and V_1 are the molarity and volume before a dilution, and M_2 and V_2 are the molarity and volume after a dilution. Consider this example from the 2006 NECTA: "Hydrochloric acid solution made by diluting 750 cm³ of 0.25 M HCl to 937.5 cm³ with distilled water."

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Keep both the residue and the filtrate. (b) To about 2 cm³ of the filtrate, add few drops of Pb (NO₃)₂ (aq)-Boil the mixture and then allow to cool. (c) (i) Put the residue in a test tube and add dilute HNO₃. Shake the mixture and divide the solution into two portions.

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