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A Answer Key

Prentice Hall Chemistry Chapter 10: Chemical Quantities... Chemical equilibrium for a reversible chemical reaction occurs when the rate of the forward reaction equals the rate of the reverse reaction. The ratio of the forward rate to the reverse rate is called the equilibrium constant.

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CHAPTER 10: Chemical Quantities

BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The

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mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

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Chapter 10: Chemical Quantities. 37 mol
of N₂ gas. 1 cup flour + 2 eggs + ½ tsp
baking powder 5 pancakes • This
relationship can be expressed
mathematically. 0 kg × 12 apples/ 1
dozen × 8 seeds/1 apple = 672 seeds
Practice Problems Plus How many
bushels of apples are in 1.

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Chapter 10 Chemical Quantities 91
SECTION 10.1 THE MOLE: A
MEASUREMENT OF MATTER (pages

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287-296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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the SI unit representing 6.02×10^{23}
representative particles of a substance.
the temperature and pressure at which
one mole of gas occupies a volume of
22.4 L. equal volumes of gases at the
same temperature and pressure contain
equal numbers of particles.

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Worksheet Chemical quantities

Determine the chemical quantities for the following, make sure your answers are in correct number of significant figures!!!

1 Calculate the number of moles for the following quantities: a 1.06×10^{23} atoms of tungsten b 3.008×10^{23} molecules of Carbon Dioxide

2

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The presentation of information in this chapter assumes that you can already perform the tasks listed below. You can test your readiness to proceed by answering the Review Questions at the end of the chapter. This might also be a

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good time to read the Chapter Objectives, which precede the Review Questions. 1.09×10^4 kg P g P mol P mol P 40 10 ...

Chapter 10 Chemical Calculations and equations

10.2 Mole to Mole & Mole to Volume Relationships *For all the problems on

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this page, first find the molar mass of the compound. 1. What is the mass of 9.45 mol of aluminum oxide? mol Al₂O₃ + 3 (1 A GO q 63,50)- 2. What is the mass of 4.52×10^{-3} mol of ethylbenzene, C₆H₅CH₂CH₃? *Ethylbenzene is a hydrocarbon that is produced by burning coal.

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Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

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Problem. A chemical bond is the

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attraction between a positive nucleus and negative electrons, or positive and negative ions. 0.2×10^{23} because that's how many atoms ...

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